Balancing Equations Activity

Name: _____

1. What number represent the Coefficient? _____
2. What number represents the Subscript? _____
3. What element is represented by the letter "S"? _____
4. How many "S's" do you have? _____
5. What element is represent by the letter "F"? _____
6. How many "F's" do you have? _____

$5\ SF_6$

Table 1:

#	Make the following Equations on your white boards	Reactants	Products	Balanced Equation
1	$H_2 + O_2 \rightarrow H_2O$			
2	$H_2O_2 \rightarrow H_20 + O_2$			
3	$Na + O_2 \rightarrow Na_2O$			
4	$N_2 + H_2 \rightarrow NH_3$			
5	$P_4 + O_2 \rightarrow P_4O_{10}$			
6	$Fe + H_2O \rightarrow Fe_3O_4 + H_2$			
7	$C + H_2 \rightarrow CH_4$			
8	$Na_2SO_4 + CaCl_2 \rightarrow CaSO_4 + NaCl$			
9	$C_2H_6 + O_2 \rightarrow CO_2 + H_2O$			
10	$Al_2O_3 \rightarrow Al + O_2$			

Analysis/Results:

1. What does "-->" mean?

2. What side of the equation are the reactants found? products?

3. Why must all chemical equations be balanced?

4. Why can't the subscripts be changed?

5. What does it mean to "simplify" the equation?

Balance the following Reactions:

1. ____ $AlBr_3$ + ____ K → ____ KBr + ____ Al
2. ____ LiCl + ____ Br_2 → ____ LiBr + ____ Cl_2
3. ____ Mn + ____ HI → ____ H_2 + ____ MnI_3
4. ____ P_4 + ____ Br_2 → ____ PBr_3
5. ____ Na_3P + ____ CaF_2 → ____ NaF + ____ Ca_3P_2